

Name \_\_\_\_\_

## Chemical Notation and Balanced Equations

**CHEMICAL NOTATION: Complete the following table**

Chemical Notation	Number of Molecules	Number of each type of atom per one molecule
$H_2O$		
$H_2SO_4$		
$CaCO_3$		
$C_6H_{12}O_6$		
$Ca_3(PO_4)_2$		
$Al_2(SO_4)_3$		
	Number of Molecules	Total number of each type of atoms present
$3H_2O$		
$5CaCO_3$		
$3Al_2(SO_4)_3$		

## TYPES OF REACTIONS

Identify the general type of reaction represented by each equation as decomposition, combination or replacement.



**Recall that there was more than one term (synonym) to describe each of the basic reaction types we studied (decomposition, combination & replacement).**

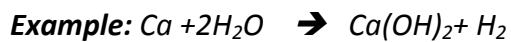
8. What are all the terms we used for a combination reaction?

9. What are all of the terms we used for a decomposition reaction?

10. What are all the terms we used for a replacement reaction?

## CHEMICAL REACTIONS: Balanced or Not Balanced?

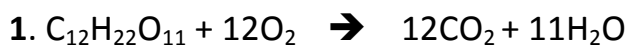
Show whether each of the following chemical reaction is balanced or not. Show your work in the same format at the example below.



1 - Ca - 1

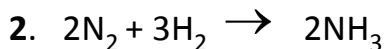
2 - O - 2     **BALANCED!**

4 - H - 4



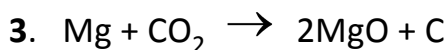
**BALANCED** or **NOT BALANCED**

(circle one and show your work)



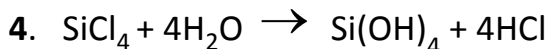
**BALANCED** or **NOT BALANCED**

(circle one and show your work)



**BALANCED** or **NOT BALANCED**

(circle one and show your work)



**BALANCED** or **NOT BALANCED**

(circle one and show your work)